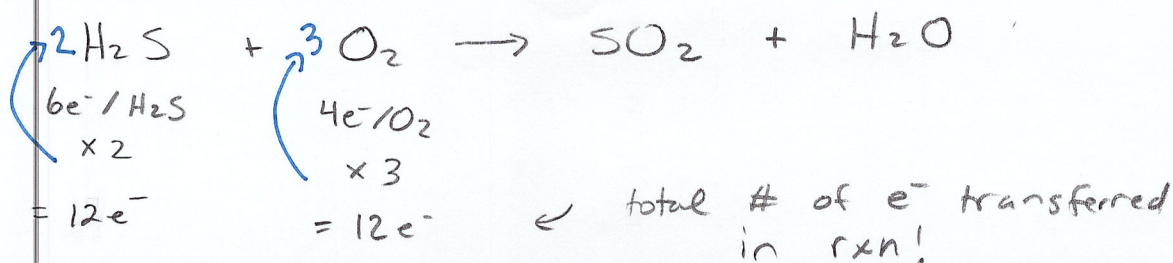
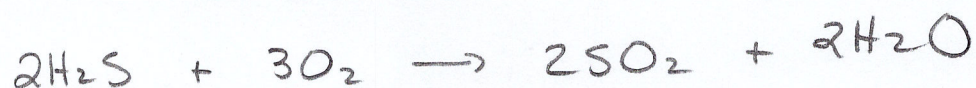


Step 4: Balance the total number of electrons transferred by finding the simplest whole number coefficient.



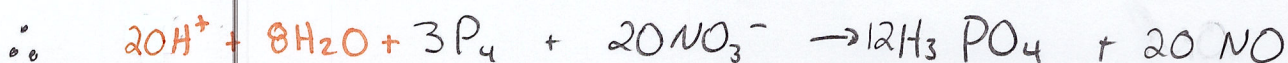
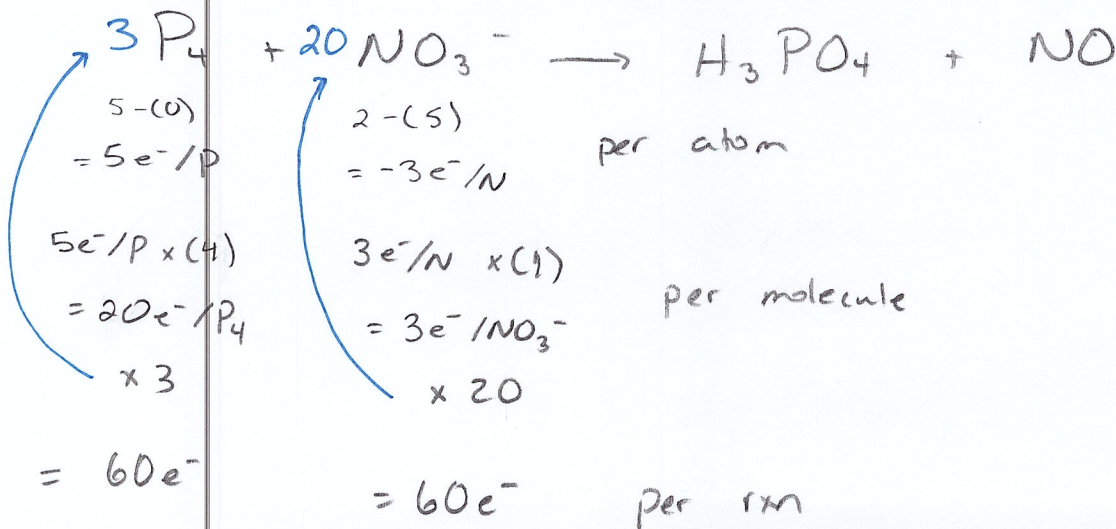
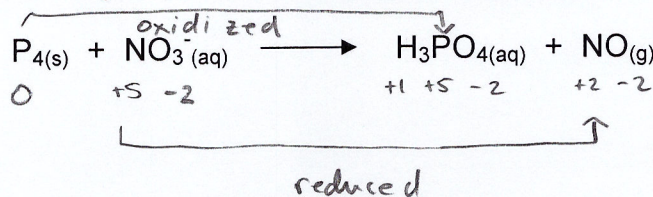
Step 5: Finish balancing the equation by balancing the number of atoms on the product side by inspection.



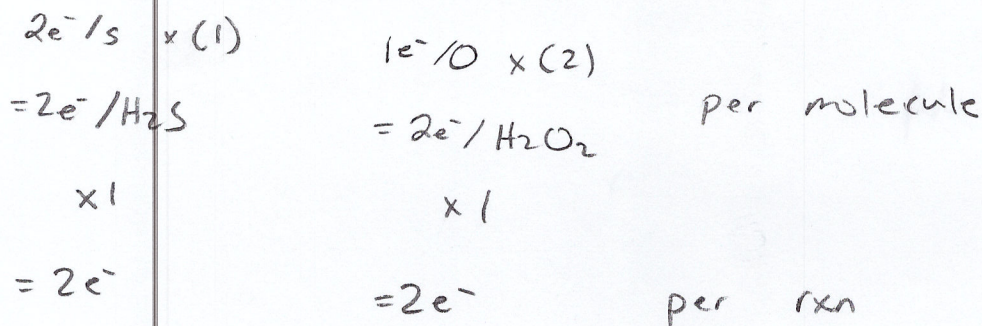
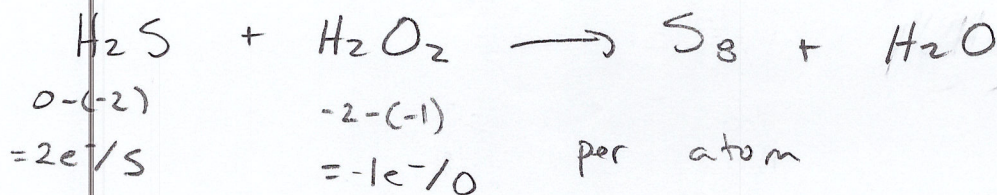
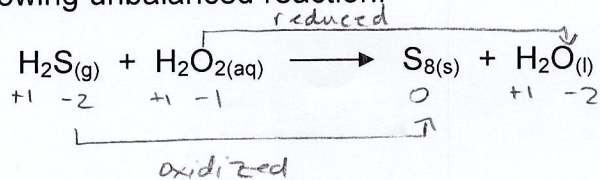
Step 6: For reactions that occur in acidic or basic solutions, include water molecules, and/or hydrogen ions as needed to balance the equation.

EXAMPLE:

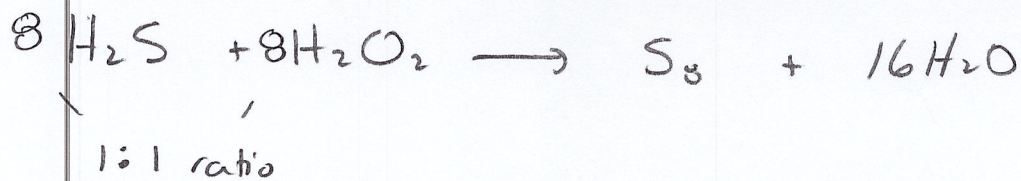
1. Balance the following redox reaction in acidic conditions:



2. Balance the following unbalanced reaction.



just needs to be in a 1:1 ratio for the coefficients of H_2S & H_2O_2



Now try pg. 464 # 15, 16a, 17(omit c)