

## NAMING IONIC COMPOUNDS

The IUPAC: International Union of Pure and Applied Chemistry, was founded in 1919 and has developed a method to

name chemicals based on their composition. The *systematic* name of a compound allows you to write the chemical's formula and predict some of its properties.

### Part A: Binary Ionic Compounds

1<sup>st</sup> name: use full name of cation first (usually metal)

2<sup>nd</sup> name: put the name of the anion last and change ending to -ide (usually non-metal)

• Name the Following Ionic Compounds:

1. NaCl (s) sodium chloride
2. LiBr (s) lithium bromide
3. CaO (s) calcium oxide
4. SrCl<sub>2</sub> (s) strontium chloride
5. BaF<sub>2</sub> (s) barium fluoride
6. K<sub>2</sub>S (s) potassium sulfide

\* why (s) → solid at room temperature

ALL ionic compounds are solid at room temp.